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Equal weight of ch4 and h2

FAQs about Matters: Gas and Liquids Hello, First of all it calculates the total number of moles using the ideal gas law. Here as you didn't mention the temperature I took it as the temperature of the room. Of formula $n = pv / rt$ you will get total moles = 0.5 Now calculates the total mass by the talque formula = moles of Ch4 + moles of so I.E $0.5 = x / 16 + x / 48$ With this, you will receive $x = 0.5$ g Now calculates the number of piers of individual formula $n = 0.5$ g / 16 for CH4 and $0.5 / 48$ for so. From the values obtained by Ch4 moles and therefore, calculates the fraction of the mole which is obtained as 0.75 for CH4 and 0.25 for so. After that, We have gas, pressure fraction = fraction of mole I.E $P1 / 12 = 0.75$ for CH4 and $12 - P1$ for so The values obtained for partial pressure for CH4 is 9 ATM and so it is 3 atm Good luck!!!! MOLAR REPORT OF CLASS 11-science A »Chemistry of CH4: H2 = 16: 2 = 8: 1 (mixed in equal quantities) The pressure is directly proportional to the number of moles. Methane is 8/9 of the total no. of mole. Thus, chA pressure fraction, a, ~

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