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TERMS

HCl
ACIDS – sour in taste, turn blue litmus red, and dissolves in water to release (H⁺) ions.
 e.g. – HCl, H₂SO₄

NaOH
BASES – bitter in taste, have soapy touch, turns red litmus blue and give (OH⁻) ions.
 e.g. – NaOH, KOH

NaCl
SALTS – it is formed by neutralization reaction between an acid and base.
 e.g. – MgSO₄

INDICATORS – it indicate the acidic or basic nature of the solution by their color change.

(a) Acid **(b) Base** **(c) Salt**



10TH CHEMISTRY NOTES

ACIDS, BASES AND SALTS

1. Acids: Acids are sour in taste, turn blue litmus red and give H₃O⁺ ion (hydronium ions in solution. e.g. HCl, H₂SO₄, HNO₃ etc

2. Bases: Bases are bitter in taste, have soapy touch, turns red litmus blue and give hydroxide ions (OH⁻) in solution. Example – NaOH, KOH etc

3. Salts: - A salt is a compound which is formed by neutralization reaction between an acid and base. For example, sodium chloride is formed by reaction between hydrochloric acid and sodium hydroxide.



4. Indicators - Indicators are substances which indicate the acidic or basic nature of the solution by their colour change. The colour of some acid-base indicators in acidic and basic medium are given below

INDICATORS	COLOUR IN ACIDIC MEDIUM	COLOUR IN BASIC MEDIUM
1. Litmus Solution	Red	Blue
2. Methyl Orange	Pink	Orange
3. Phenolphthalein	Colourless	Pink
4. Methyl Red	Yellow	Red

5 Chemical properties of acids

(i) Acids react with active metals to give hydrogen gas.



(ii) Acids react with metal carbonate and metals hydrogen carbonate to give carbon dioxide.



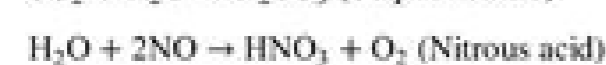
(iii) Acids react with bases to give salt and water. This reaction is called neutralization reaction.



(iv) Acids react with metals oxides to give salt and water.

Acid	Base	Salt	Example
Strong	Strong	Neutral	NaOH + HCl → NaCl + H ₂ O
Strong	Weak	Acidic	HCl + NH ₄ OH → NH ₄ Cl + H ₂ O
Weak	Strong	Basic	CH ₃ COOH + NaOH → CH ₃ COONa + H ₂ O
Weak	Weak	Neutral	CH ₃ COOH + NH ₄ OH → CH ₃ COONH ₄ + H ₂ O

impure form of water, contains CO_2 , SO_2 , NO . These oxide gases react with water and get chemically converted into acids which are responsible for the release of ions that conduct electricity—contains many ionic species (acids) whose pH is less than 7 and therefore it conducts electricity



8. Why do acids not show acidic behaviour in the absence of water?

Solution:

A substance is said to be acidic when it can generate hydrogen ions (H^+). The hydrogen ion generally comes from the acid which dissociates in the presence of water. Thus, for acid to dissociate into hydrogen ion and the respective anion, water must be present. Hence, an acid fails to show its acidic behaviour in the absence of water.

9. Five solutions A, B, C, D, and E, when tested with universal indicator, showed pH as 4, 1, 11, 7 and 9, respectively. Which solution is

- Neutral?
- Strongly alkaline?
- Strongly acidic?
- Weakly acidic?
- Weakly alkaline?

Arrange the pH in increasing order of hydrogen-ion concentration.

Solution:

At 25°C pH of neutral solutions = 7

As the pH falls below 7, it denotes acidic character with a pH of 1 being highly acidic. When the pH goes above 7, it implies that the solution is basic with 14 being highly basic.

- Neutral Solution D with pH 7
- Strongly alkaline Solution C with pH 11
- Strongly acidic Solution B with pH 1
- Weakly acidic Solution A with pH 4
- Weakly alkaline Solution E with pH 9

What is formed when zinc reacts with sodium hydroxide? The base is the base, the more above is the pH. Organic chemistry (5th ed.). However, Lewis suggested that an electron pair donor was classified as a base and an electron pair of pair was classified as acid. Sodium carbonate is a basic salt because it is a salt of (a) strong acid and a strong base (b) weak acid and weak base (c) strong acid and weak base (d) weak acid and response of strong base: D 12. The role of the calcium chloride assumed in the protection tube is to (a) absorb the advanced gas (b) moisten the gas (c) absorb the humidity from the gas (d) absorb the advanced gas (e) absorb the advanced gas: C explanation: Reason: Dries of the guard tube (absorbs water) from calcium chloride on a humid day. See also Basic Acid (Chemistry) Acid "Basic Reaction BRUATA É i É á,—" Lowry acid "theory of the Base Chiral Lewis frustrated acid lewis pair Gutmann" Beckett Method ECW References ^ a b iupac, compendium of chemical terminology, 2nd ed. ^ A b March, J. which of the following statements is correct on an aqueous solution of an acid and a base? In 1923, Lewis wrote an acidic substance is the one that can use a solitary electronic couple from another molecule in completing the stable group of one of its atoms. [2] [16] The theory of the Base of BRAVA "Lowry acid" was published in the same year. (1996). As usual, a weak acid has a strong conjugated base. (i) The bulb does not shine because the electrolyte is not acidic. If you have any questions regarding the science of the CBSE 10 chapter 2 Acids bases and multiple choice questions with the answers, leave a comment in a row and we will return to you soon. BIBCODE: 1977]ched.54.612c. Many metallic complexes act as Lewis acids, but usually only after dissociating a more weakly tied Lewis base, often in another confrontation between Lewis and Brá É i É ávelop "Lowry aciditate by Brown and Kanter. [18] 2,6-of-Butylpyridine reacts to form the il Salt with HCL but does not react with BF3. "On the nature of the dative bond: coordination with metals and beyond. Lewis. [2] The nucleophilic and electrophil terms are more or less interchangeable with the Lewis base and Lewis acid, respectively. Composed of O, S, if e and You in the Oxidation State -2, including water, Ethers, ketones, the foundations of Lewis more common are anions. Doi: 10.1351/Goldbook.103508 ^ a b Lewis, Gilbert Newton (1923). 11. Na2co3. Doi: 10.1021/Ja02261A002. Furthermore, in some cases (for example, sulfoxids and amine oxides such as R2S á É á É 'o and R3n á É á É 'o), the use of the arrow of the dative bond is only a notional comfort to avoid the Drawing of formal charges. (1977). (a) hydroxide of hydroxide and sodium zinc (b) sodium galvanized and hydrogen (c) sodium oxide of zinc and hydrogen (d) sodium galvanized and response/ explanation of water response: B explanation : Reason: ZN + 2NaOH á É á É á É MA2ZN02 (sodium galvanized) + H2 6. However, these terms, in particular their abstract noun, forms nucleophilia and el Elterophilia, emphasize the kinetic aspect of Reacti Vity, while Lewis' basicit has and Lewis' acidite underline the thermodynamic aspect of the formation of Lewis' decorations. [3] Which depicting the decorations in many cases, the interaction between the Lewis base and the Lewis acid in a complex is indicated towards Lewis acid using the notation of a dative bond - for example, me3bá É á É NH3. The equation involves the inversion of acids and basic strengths. The conjugated base of an acid Lowry by BRANSSted is also a base of Lewis as a loss of H+ from the acid leaves those electrons that have been used for the binding H as a solitary couple on the conjugated base. 142: "We are inclined to think about substances as the possession of acid or basic property, without having a particular solvent in mind.) Water

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